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Importance of electrolysis in our daily lives

INTRODUCTION
Electrolysis is a process that uses electrical energy to drive a non-spontaneous redox reaction. This process is essential in many industries, including metal refining, electroplating, and the production of various chemicals. The importance of electrolysis lies in its ability to produce pure metals and other valuable products that are difficult to obtain through other means. This document explores the various applications of electrolysis and its significance in modern industry.

APPLICATIONS OF ELECTROLYSIS

1. Metal Refining: Electrolysis is used to purify metals such as copper, zinc, and aluminum. Impure metal is placed in an electrolytic cell, and a more reactive metal is used as the anode. The impurities are dissolved into the electrolyte, while the pure metal is deposited at the cathode.
2. Electroplating: This process involves coating a metal object with a thin layer of another metal. The object to be plated is connected to the negative terminal of a power source, and a metal salt solution is used as the electrolyte. The metal ions in the solution are reduced at the cathode, forming a uniform coating on the object.
3. Production of Chlorine and Hydrogen: Electrolysis of brine (sodium chloride solution) is used to produce chlorine gas and hydrogen gas. Chlorine is used in the production of various chemicals, while hydrogen is used in the synthesis of ammonia and other industrial processes.
4. Production of Sodium and Potassium: Electrolysis of molten sodium chloride and potassium chloride is used to produce these highly reactive metals. These metals are used in various applications, including the production of sodium lamps and potassium permanganate.

CONCLUSION
Electrolysis is a vital process in modern industry, enabling the production of pure metals and other valuable products. Its applications range from metal refining and electroplating to the production of chlorine, hydrogen, and sodium. Understanding the principles and applications of electrolysis is essential for anyone interested in chemistry and industrial processes.

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ELECTROLYSIS PRACTICE PROBLEMS (9.5 & 19.2) ... Aqueous solutions containing different concentrations of NaCl were electrolysed using platinum electrodes.

by Z Yan · 2020 · Cited by 50 — critical problems that the world faces in this century. Pop- ... The use of buffered solutions in CO2 electrolysis creates.. Question 1. Calculate the quantity of electricity obtained from 2 moles of electrons. Solution: (Based on the StoPGoPS approach to problem solving).

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The amount of a substance deposited on each electrode in an electrolytic cell is directly ... Copper(II) sulfate solution is toxic by ingestion and is a.. [1] Ans. (a) Electrolytic dissociation : The process in which separation of ions of an ionic compound takes place on heating or in an aqueous solution is ...

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